

Synthesis, Electrochemical Characterization and Photocatalytic Application of Ce ion Doped ZnO nanoparticles using Leaf Extract of *Sesbania Grandiflora* by Green Method

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Abstract: Nanosized ZnO particles of specific morphology were synthesized using the plant leaf extracts of *Sesbania grandiflora* (*Agathikeerai*). The structures and morphology of these fabricated ZnO nanoparticles and Ce ion doped ZnO nanoparticles were characterized by FTIR, SEM-EDAX, AFM and PL. Photodegradation and electrochemical behavior of the nanoparticles were studied. The aqueous *Sesbania grandiflora* leaf extract acts as a solvent with multiple roles as promoter and capping agent for the synthesis of ZnO and Ce ion doped ZnO nanoparticles. Photocatalytic degradation was also investigated with Bismarck brown dye under UV-irradiation source. The ZnO and Ce ion doped ZnO nanoparticles exhibited potential photocatalytic activity towards the degradation of Bismarck brown dye. Green synthesis using *Sesbania grandiflora* is found to be the best capping agent for synthesizing nanoparticles.

Keywords: Zinc oxide nanoparticles, Green synthesis, *Sesbania grandiflora*, FT-IR, SEM-EDAX, AFM, PL, Photodegradation, Electrochemical studies

1. Introduction

Green synthesis techniques make use of moderately pollutant free chemicals to synthesis nanomaterials and embrace the use of benign solvents such as water, natural extracts. Green chemistry seeks to reduce pollution at source [1,2]. The advantage of using plants for the synthesis of nanoparticles is that they are easily available, safe to handle and possess a broad variability of metabolites that may aid in reduction. A number of plants are being currently investigated for their role in the synthesis of nanoparticles. ZnO belongs to the class of metal oxides, which is characterized by photo catalytic and photo-oxidising capacity against chemical and biological species [3]. Green synthesis of zinc oxide nanoparticles using Aloe vera [4], gold nanoparticles by alfalfa [5], Cinnamomum camphora [6], neem [7], Emblica officinalis [8], lemongrass [9] and tamarind [10] has been reported. Therefore the present investigation has been made to synthesize ZnO nanoparticles by *Sesbania grandiflora*.

Sesbania grandiflora

Sesbania grandiflora L. is a plant from family fabaceae cultivated in all over India for its edible flowers. It has synonym *Agathis grandiflora* (Fig.1) and commonly known as Hummingbird Tree, Butterfly Tree. It's one of the well-known medicinal plants of India. *Sesbania grandiflora* has been known to have antimicrobial activities [11, 12].



Figure 1: *Sesbania grandiflora*

It is a 3–8 m long and the colour of the flowers is red and white in colour. It contains alkaloids, phenols, flavonoids, carbohydrates, saponins, phytosterols, Triterpenes [13], amino acids and alcohol [14]. Leaves are used to disinfect throat and cure kidney diseases. Saponins have been extensively used as detergents, pesticides and molluscicides and also have beneficial health effects.

2. Experimental Methods

2.1 Collection of Plant

The plant *Sesbania grandiflora* was collected in the surroundings of Tirunelveli district, Tamilnadu.

2.2 Preparation of the leaf extract

The collected *Sesbania grandiflora* leaves were washed several times with water to remove the impurities. Leaves had been dried in the sun shade for 7 days. After the leaves were dried, it is powdered using mortar. The *Sesbania grandiflora* leaf powder of 10g was used for

synthesis purpose. The weighed 10g leaf powder was mixed gently with 100ml of distilled water, boiled in 80°C for 60min; until the colour of the solution changes from watery to yellow. Then the extract was filtered through Whatman No.1 filter paper, stored at room temperature in order to use for further studies.

2.3 Synthesis of ZnO nanoparticles and Ce ion doped ZnO nanoparticles using *Sesbania grandiflora* leaf extract:

5g Zinc Nitrate was added to the leaf extract under vigorous stirring. After 15 min stirring the solution was heated at a temperature of about 60°C. The yellow coloured paste was finally calcined at 400°C for 2h. The ZnO nanoparticles so obtained were preserved in the air-tight vials for further studies [15]. Similar procedure was adopted for the synthesis of Ce ion doped ZnO nanoparticles using 1% cerium nitrate as the precursor material.

2.4 Photocatalytic Measurement:

Nanosized ZnO particle is a good photocatalyst to degrade organic contaminants, such as Bismarck brown dye. The dye solution was prepared by dissolving 10mg powder of Bismarck brown dye in 100ml distilled water. 0.1g nanoparticles was added to 100ml of prepared Bismarck brown dye solution and the mixer was stirred magnetically for 60min in shadow before exposing to sunlight. Then the colloidal suspension was placed in a closed chamber and irradiated with sunlight. The reactions were observed one by one in every time interval of 10 min for 1hr. Finally, the rate of dye decomposition was monitored by taking 4ml samples from each set and recording the UV-Vis spectra in the wavelength after centrifugation and filtration [16].

2.5. Instrumentation

Nicolet iS5 instrument was used to identify the functional groups of the synthesized zinc oxide nanoparticles. The surface morphology of the nanoparticles was done using Carl Zeiss EVO 18 SEM operating at 15 KV using normal incidence. EDAX measurements were carried out by Quantax 200 with X-Flash-Bruker. The electrochemical analyzer, CH Instruments Electrochemical Workstation model 650C was employed for various electrochemical studies performed in this study.

3. Results and Discussion

3.1. FT-IR Studies:

The FTIR spectrum of ZnO nanoparticles synthesized using aqueous leaf extract of *Sesbania grandiflora* was shown in the Fig.2. The broad peak located at 3444.88 cm⁻¹ can be

assigned to the O-H stretching vibrations, indicating the presence of hydroxyl groups [17]. Few less intense peaks centered at 2923.69 cm⁻¹, 2853.34 cm⁻¹ and 2426.27 cm⁻¹ are probably due to presence of aliphatic asymmetric C-H stretching vibration, C-H stretching and O-H stretching in carboxylic acid respectively [18]. The peak at 1795.46 cm⁻¹ can be assigned to the -C=O- group and the peak at 1763.17 cm⁻¹ can be assigned to the C-O-C stretching of polysaccharides present in the *Sesbania grandiflora* leaf extract [19]. Moreover, the peaks at 1384.24 cm⁻¹ and 1627.76 cm⁻¹ were mainly attributed to the -C=N- stretching vibrations as well as amide bands of proteins [20]. In addition, there is a peak at 1421.91 cm⁻¹ corresponds to O-H bending of the hydroxyl group present. The bands at 1162.65 and 1121.27 cm⁻¹ can be attributed to the -C-O- stretching and O-H deformation assigned to the water molecule respectively.

The frequency observed at 917.52 cm⁻¹ corresponds to O-H bending of carboxylic acid present. The peak at 875.25 cm⁻¹ was due to the formation of tetrahedral coordination of Zn ion [21]. The peaks observed at 825.62 cm⁻¹ and 839.32 cm⁻¹ are due to the weak band (result of C-H out of plane bending) and aromatic -C-H- out of plane bending respectively. The peak at 712.87 cm⁻¹ is due to the presence of R-CH group. The peak at 679.40 cm⁻¹ indicates the stretching vibrations of ZnO nanoparticles which is consistent with the reported data [22]. The characteristic peak appeared at 439.38 cm⁻¹ could be attributed to the metal oxygen (Zn-O) bond.

Fig.3. shows the FTIR spectrum of Ce ion doped ZnO nanoparticles synthesized using aqueous leaf extract of *Sesbania grandiflora*. The broad peak located at 3444.51 cm⁻¹ can be assigned to the O-H stretching vibrations, indicating the presence of hydroxyl groups. The peak centered at 2925.62 cm⁻¹ corresponds to aliphatic asymmetric -C-H- stretching vibration. The peaks at 1500.26 cm⁻¹ and 1629.41 cm⁻¹ are due to -C-H- stretching and -C=O- stretch in polyphenols respectively. The bands at 1384.20 cm⁻¹ and 1141.35 cm⁻¹ can be attributed to the aromatic -C=C- bond, -O-C- stretching respectively. The peak located at around 1063.58 cm⁻¹ can be assigned as the absorption peak of -C-O-C- or -C-O- stretching. The peak around at 596.39 cm⁻¹ indicating the formation of stretching mode of Ce ion doped ZnO [23]. The peaks corresponding to Zn-O bonds are shifted towards lower wavenumber for Ce doped ZnO nanoparticles, indicating the incorporation of Ce ions in the ZnO lattice [24]. The bands obtained at 659.53, 612.25, 496.91 and 443.59 cm⁻¹ are due to the characteristic stretching vibrations of ZnO nanoparticles [25].

3.2. FT-IR Spectrum

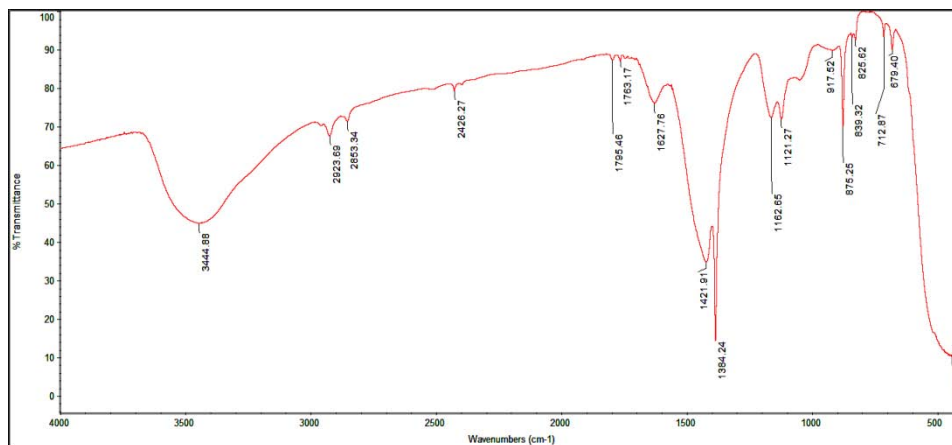


Figure 2: FTIR spectrum of ZnO nanoparticles synthesized using *SesbaniaGrandiflora* leaf extract

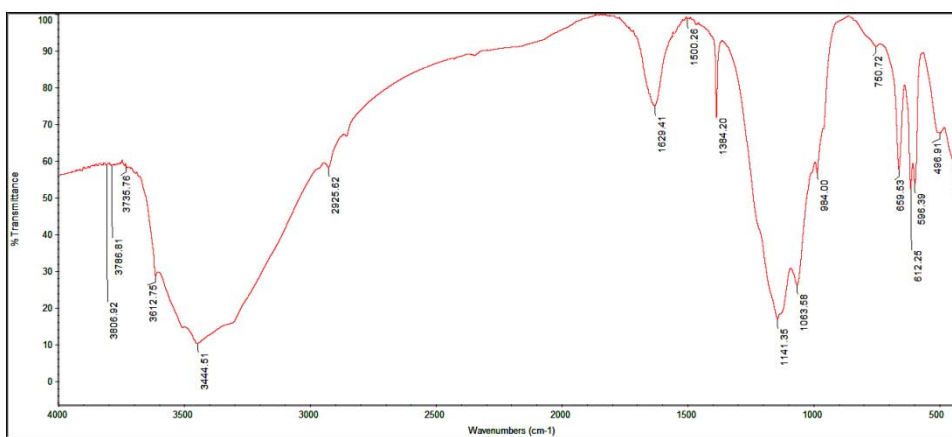


Figure 3: FTIR spectrum of Ce ion doped ZnO nanoparticles synthesized using *SesbaniaGrandiflora* leaf extract

3.3. SEM Analysis

Fig.4 represents the SEM image of ZnO nanoparticles synthesized using *SesbaniaGrandiflora* leaf extract. This picture substantiates the approximate spherical shape to the ZnO nanoparticles with a granular nature, and most of the particles exhibit some agglomeration [26]. The aggregation of particles should have been originated from the large specific surface area and high surface energy of ZnO nanoparticles [27]. The aggregation occurred may be probably due to the drying process [28, 29]. Fig.5 shows the SEM image of Ce ion doped zinc oxidenanoparticles synthesized using *SesbaniaGrandiflora* leaf extract and it exhibited distinct spherically balllike structure.

3.4. SEM Image

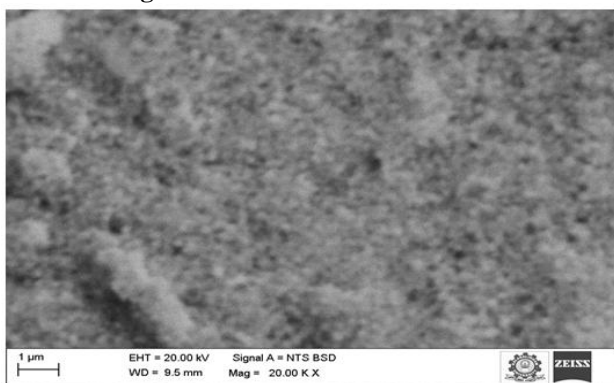


Figure 4: SEM image of ZnO nanoparticles synthesized using *Sesbania Grandiflora* leaf extract

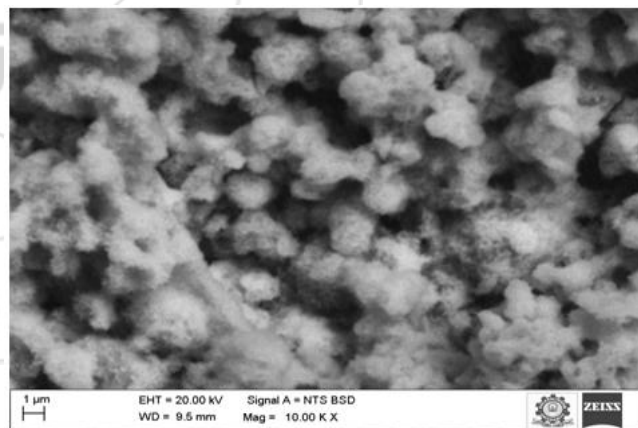


Figure 5: SEM image of Ce ion doped ZnO nanoparticles synthesized using *Sesbania Grandiflora* leaf extract

3.5. Energy Dispersive X-Ray Analysis

In order to confirm the presence of Ce ion doped ZnO nanoparticles synthesized using *SesbaniaGrandiflora* we perform the EDAX spectroscopy. Fig.6. shows the EDAX spectrum of Ce ion doped zinc oxide nanoparticles synthesized using *SesbaniaGrandiflora* leaf extract. Ce ion doped zinc oxide nanoparticles were found to have atomic percentage 37.61 of Zn, 19.86 of O, 6.31 of Ce as shown in Table 1.This confirmed the doping of Ce ion in ZnO lattice.

3.6. EDAX Image

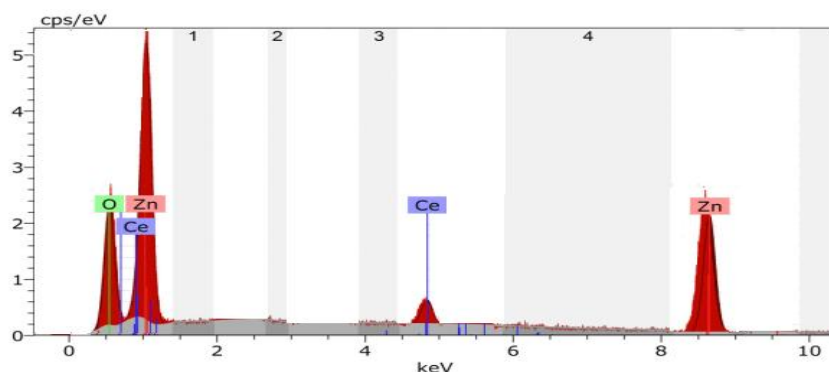


Figure 6: EDAX spectrum of Ce ion doped ZnO nanoparticles synthesized using *Sesbaniagrandidflora* leaf extract

Table 1: Atomic composition of Ce ion doped ZnO nanoparticles synthesized using *Sesbaniagrandidflora* leaf extract

Element	Series	Unn.C [wt.%]	Norm.C [wt.%]	Atom.C [at.%]	Error (3 sigma) [wt.%]
Zinc	K-series	37.61	58.97	30.90	3.22
Oxygen	K-series	19.86	31.13	66.68	7.68
Cerium	L-series	6.31	9.90	2.42	0.64
Total		63.78	100.00	100.00	

the AFM spectral image of ZnO nanoparticles synthesized using *Sesbaniagrandidflora* leaf extract with a scanning area to 0m X 3.13µm, we found spongy shape distributed over the surface [30]. These particles are between 0m Y 3.13µm in length and the size is in the range of 20-50nm. Fig.8.shows the AFM spectral image of Ce ion doped ZnO nanoparticles synthesized using *Sesbaniagrandidflora* leaf extract with a scanning area to 952nm X 2.19µm. These particles are between 1.89µm Y 3.13µm in length. The shape of Ce ion doped ZnO nanoparticles is spongy layered form and the size is in the range of 30-50nm.

3.7. AFM Spectral Studies

AFM spectra were recorded for the ZnO and Ce ion doped ZnO nanoparticles deposited on the glass plate. Fig.7. shows

3.8. AFM Image

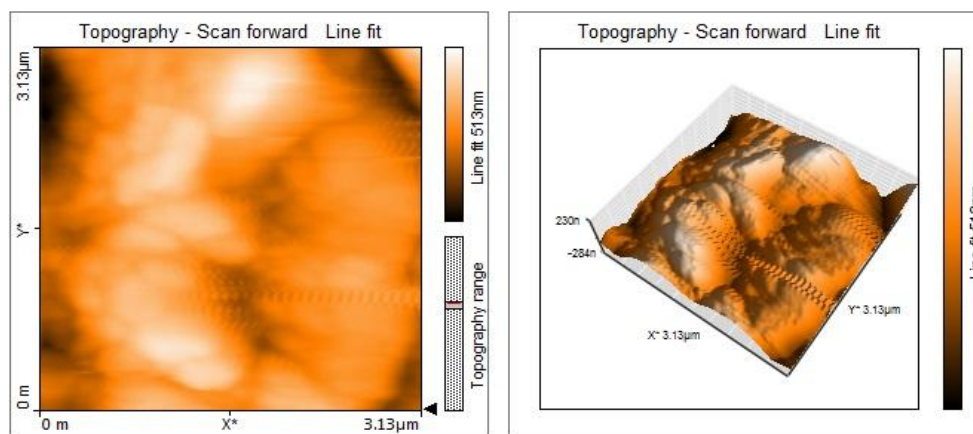


Figure 7: AFM image of ZnO Nanoparticles synthesized using *SesbaniaGrandiflora* leaf extract

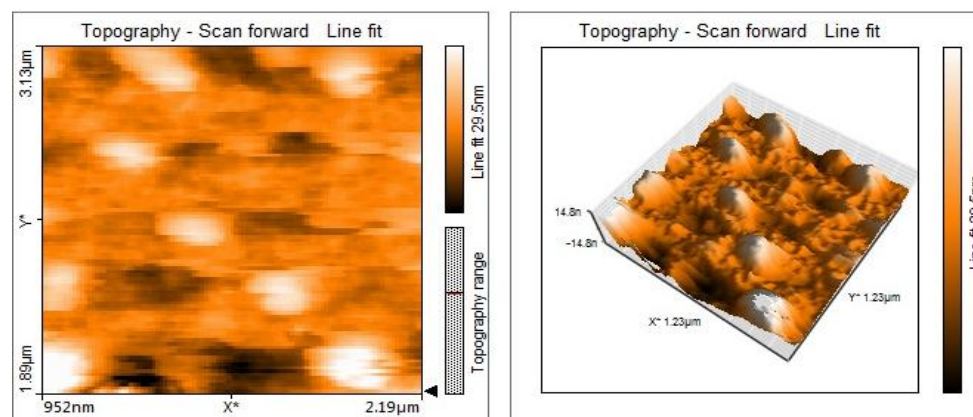


Figure 8: AFM image of Ce ion doped ZnO Nanoparticles synthesized using *SesbaniaGrandiflora* leaf extract

3.9. Photoluminescence

PL spectra were measured for the ZnO and Ce ion doped ZnO nanoparticles synthesized using aqueous leaf extract of *Sesbaniagrاندiflora* in the range of 200-700 nm are shown in Fig.9 and Fig.10. The wavelength of excitation chosen for the samples is 350 nm. It is interesting to note from that although the excitation spectra of ZnO and Ce ion doped ZnO nanoparticles appear similar, an appreciable change in luminescence intensity is observed in the doped samples compared to the undoped samples. For doped samples there is increase in excitation intensity. This might be due to the interaction between Ce ion and ZnO lattice. This provides an obvious evidence for the entry of Ce ion in the ZnO lattice. Photoluminescence investigation evidenced the high crystalline nature of the undoped and doped ZnO nanoparticles.

3.10. PL Spectrum

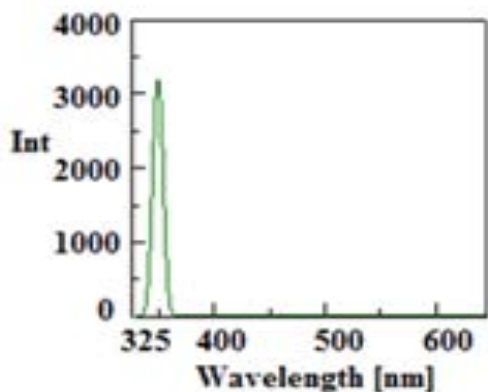


Figure 9: PL spectra of ZnO nanoparticles

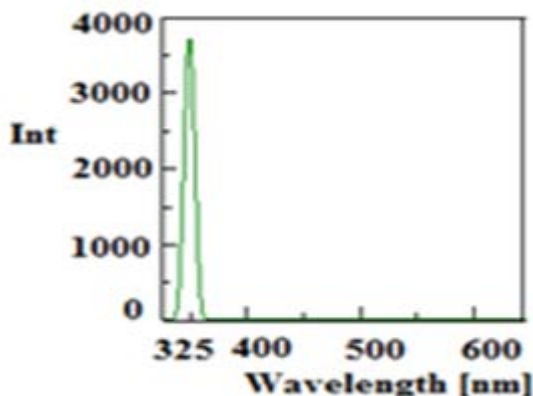


Figure 10: PL spectra of Ce doped ZnO nanoparticles

3.11 Electrochemical Analysis

3.11.1 Electrochemical Impedance Measurements

EIS measurements were obtained in the frequency range from 100000Hz to 0.01Hz by using amplitude of 0.5V. EIS was applied to study the resistance ability of the ZnO and Ce ion doped ZnO nanoparticles. The Nyquist plot representation of impedance spectra of green synthesized ZnO and Ce ion doped ZnO nanoparticles in pH=1 medium were shown in Fig.11 and Fig.12. The different parameters were tabulated (Table.2). The values of charge transfer resistance increases for ZnO and Ce ion doped ZnO nanoparticles when compared to the bare. The values of conductance are high for ZnO and Ce ion doped ZnO

nanoparticles when compared to the bare. These facts reveal that both the ZnO and Ce ion doped ZnO nanoparticles have high conductivity.

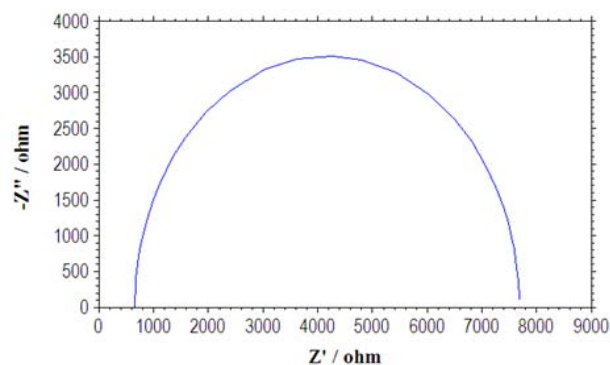


Figure 11: EIS of undoped ZnO nanoparticles synthesized using *SesbaniaGrandiflora* Leaf extract

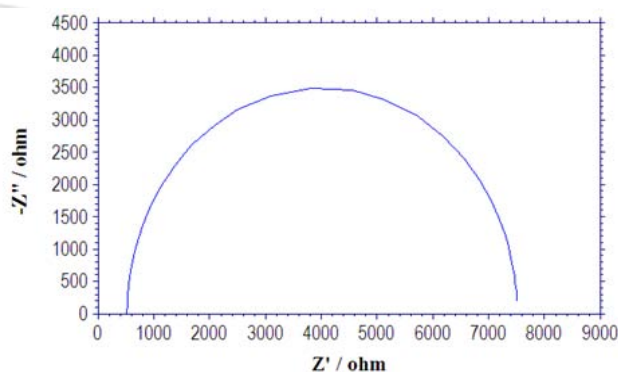


Figure 12: EIS of Ce ion doped ZnO nanoparticles synthesized using *SesbaniaGrandiflora* Leaf extract

Table 2

Name of Plant material	Nanoparticles	f_{max}	R_{ct} (Ωcm^{-2})	C_{dl} ($\mu F cm^{-2}$)
	Bare	2165	4314	1.704
<i>SesbaniaGrandiflora</i>	ZnO	3559	6922	6.46
	Ce:ZnO	3546	6976	6.43

3.11.2. Cyclic Voltammetry

Cyclic voltammograms were recorded in the pH 1.0 for 1.0 ml of ZnO and Ce ion doped ZnO nanoparticles synthesized using aqueous leaf extract of *SesbaniaGrandiflora*. The Glassy Carbon Electrode was (GCE) as working electrode Vs Ag/AgCl. ZnO nanoparticles and Ce ion doped ZnO nanoparticles showed one oxidation peak and one reduction peak in pH 1 conditions. The background current was recorded for all sweep rates studied in the potential range from -1.0 to 1.0 V and subtracted properly in calculating the peak currents.

The cyclic voltammetric behaviour of 1.0 ml of ZnO nanoparticles in 0.1M H_2SO_4 (pH 1.0) was studied at GCE as shown in Fig.13. In this cyclic voltammogram, one sharp anodic peak and one cathodic peak were observed in the potential range from -1.0V to 1.0V. The cathodic and anodic peaks were seen around the potential at -0.7123 V and 0.3256 V respectively. The values of peak current and peak potential are presented in Table 3.

The cyclic voltammetric behaviour of 1.0 ml of Ce ion doped ZnO nanoparticles in 0.1M H₂SO₄ (pH 1.0) was studied at GCE as shown in Fig.14. In this cyclic voltammogram, one sharp anodic peak and one small cathodic peak were observed in the potential range from -1.0V to 1.0V. The cathodic and anodic peaks were seen around the potential at -0.6789V and 0.4291V respectively. The values of peak current and peak potential are presented in Table 3. The peak potentials obtained for doped samples were entirely different from the behaviour obtained for undoped samples confirming the formation of doped samples.

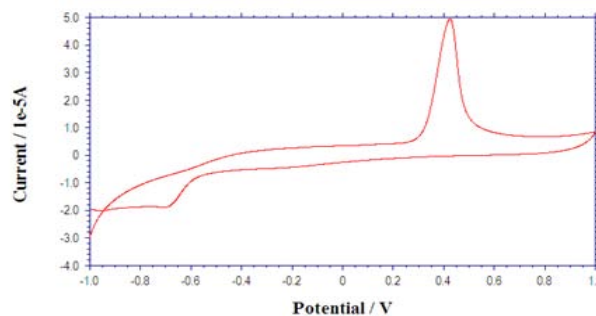


Figure 14: Cyclic Voltammogram of Ce ion doped ZnO nanoparticles synthesized using *Sesbania Grandiflora* leaf extract

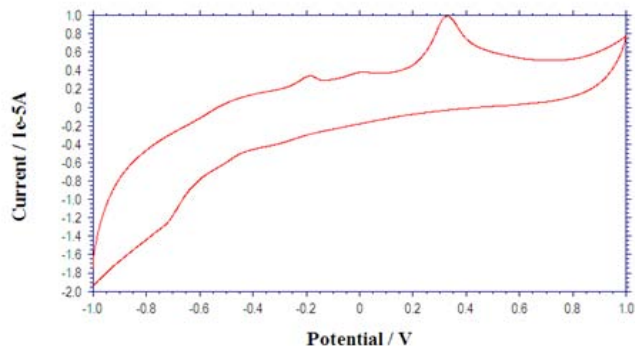


Figure 13: Cyclic Voltammogram of ZnO nanoparticles synthesized Using *Sesbania Grandiflora* leaf extract

Table 3: Cyclic voltammetric behaviour data of ZnO and Ce ion doped ZnO nanoparticles in pH 1.0

Name of Plant material	Nanoparticles	Oxidation		Reduction	
		E(V)	I(A)	E(V)	I(A)
<i>Sesbania Grandiflora</i>	ZnO	0.3256	9.664×10 ⁻⁶	-0.7123	-1.247×10 ⁻⁵
	Ce:ZnO	0.4291	4.899×10 ⁻⁵	-0.6789	-1.839×10 ⁻⁵

3.12. Photocatalytic Activity

The UV visible absorbance values of pure Bismarck brown dye solution shows absorption wavelength at 460nm. The characteristic absorbance value at 460nm was used to track the photocatalytic degradation process. Fig.15 can be clearly noticed from the recorded values that no significant changes of the concentration of Bismarck brown dye after 3 hrs irradiation, which indicated that pure Bismarck brown dye solution, cannot be easily degraded by UV light. The degradation efficiency of pure Bismarck brown dye within 3 hrs irradiation time was about 28%. The result showed that the photocatalytic activity of pure Bismarck brown dye was very less when compared with the ZnO nanoparticles and Ce ion doped ZnO nanoparticles synthesized using *Sesbania grandiflora* leaf extract. The dye degradation in presence of bio synthesized nanoparticles was verified by the decrease of the peak intensity during 60min exposure in solar light shown in Fig.16 and Fig.17. The dye degradation (%) was calculated by using the following equation (2).

$$\text{Dye degradation (\%)} = \frac{C_0 - C_t}{C_0} \times 100\% \quad (1)$$

Where C_0 is the initial concentration of Bismarck brown and C_t is the concentration of the dye solution at the selected irradiation time. Fig.18. shows the bleaching of Bismarck brown dye on photodegradation in the presence of ZnO and Ce ion doped ZnO nanoparticles as photocatalyst. The degradation efficiency was higher in the presence of Ce ion doped ZnO nanoparticles than that for ZnO nanoparticles.

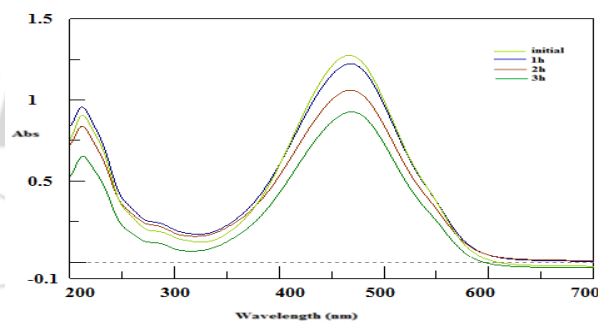


Figure 15: Photocatalytic dye degradation of pure Bismarck brown

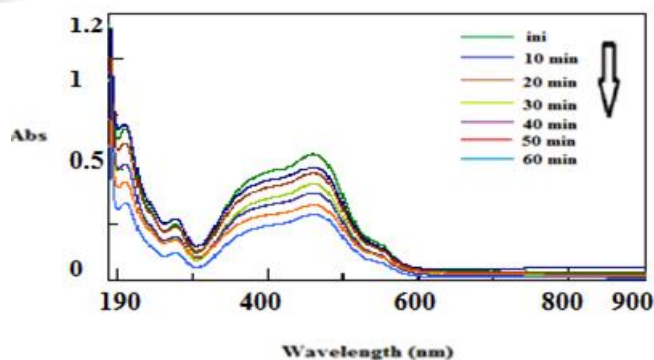


Figure 16: Photocatalytic dye degradation of Bismarck brown in presence of ZnO nanoparticles synthesized using *Sesbania grandiflora* leaf extract

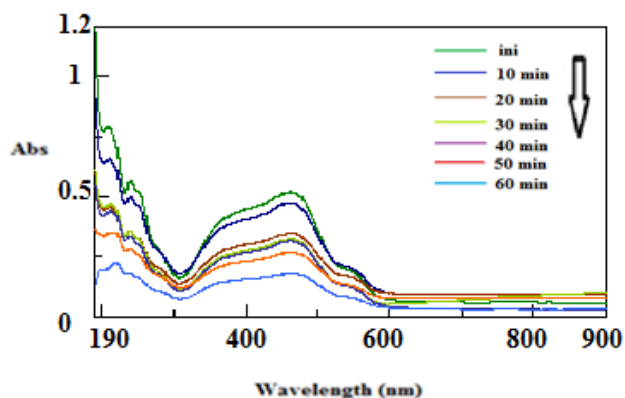


Figure 17: Photocatalytic dye degradation of Bismarck brownin presence of Ce ion doped ZnO nanoparticlessynthesized using *Sesbaniagrandidiflora* leaf extract

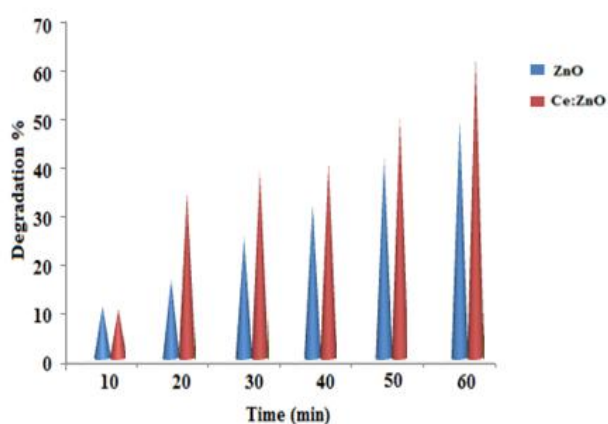


Figure 18: Degradation of Bismarck browndye in the presence of ZnO and Ce ion doped ZnO nanoparticles synthesized using *Sesbaniagrandidiflora* leaf extract

3.13 Conclusion

ZnO and Ce ion doped ZnO nanoparticles are synthesized (simple and cost effective) using aqueous leaf extract of *Sesbaniagrandidiflora* (Agathikeerai). The FT-IR studies showed an absorption peak at 437cm^{-1} (Zn-O linkage) which indicated the formation of zinc oxide nanoparticles. The aqueous *Sesbaniagrandidiflora* leaf extract reveals the presence of phyto constituents like alcohol, aldehyde and amine which were the surface active molecules, stabilize the nanoparticles. The surface morphology of the ZnO and Ce ion doped ZnO nanoparticles was characterized by SEM analysis and suggested different morphological structures. Chemical purity and stoichiometry of the samples were investigated by EDAX Spectroscopy, in order to confirm the presence of Zn, O and Ce ions in the material. AFM study reveals the surface morphology of the synthesized ZnO and Ce ion doped ZnO nanoparticles. PL Spectra revealed the change of luminescence intensity in the Ce ion doped sample compared to the ZnO sample with almost a linear increase in excitation intensity. Impedance spectroscopy is employed to study the conductivity of nanoparticles. The conductivity of chemically synthesized nanoparticles is high and is used as an electronic material. The Ce ion doped ZnO nanoparticles have high R_{ct} and low C_{dl} values show effective resistance. From cyclic voltammetric studies, good

redox behaviour was observed for ZnO and Ce ion doped ZnO nanoparticles synthesized using *Sesbaniagrandidiflora*. The photocatalytic study concludes that these bio-ZnO nanoparticles have efficiency to degrade Bismarck brown dye under solar irradiation. Therefore they can find application in water purification and textile industries.

4. Acknowledgement

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References

- [1] Tundo P. and Anastas P., Eds, Green Chemistry: Challenging Perspectives. *Oxford University Press*, Oxford, UK, (2000).
- [2] Reed S.M. and Hutchison J.E., "Green Chemistry in the organic teaching laboratory: an environmentally benign synthesis of adipic acid", *J Chem Educ* 77, (2000)1627-1628.
- [3] Deepali Sharma, J.R., B.S. Kaith, Mohinder Kaur, Sapna Sharma, "Synthesis of ZnO nanoparticles and study of their antibacterial and antifungal properties", *Thin Solid Films* 519, (2010)1224-1229.
- [4] Sangeetha, G., Rajeshwari, S. and Venkatesh, R. Green synthesis of zinc oxide nanoparticles by aloe barbadensis miller leaf extract: Structure and Optical properties, *Materials Research Bulletin.*, 46(12), (2011) 2560–2566.
- [5] Gardea-Torresdey, J.L., Parsons, J.G., Gomez, E. and Peralta Videal, J. Formation and growth of Au nanoparticles inside live alfalfa plants, *Nanoletters.*, 2(4), (2002) 397–401.
- [6] Huang, J., Li, Q., Sun, D., Lu, Y., Su, Y., and Yang, X. Biosynthesis of silver and gold nanoparticles by novel sundried *Cinnamomum camphora* leaf, *Nanotechnology.*, 18(10), (2007) 105104–105114.
- [7] Shiv Shankar, S., Rai, A., Ahmad, A., and Sastry, M. Rapid synthesis of Au, Ag, and bimetallic Au core–Ag shell nanoparticles using neem (*Azadirachta indica*) leaf broth, *Journal of Colloid and Interface Science.*, 275(2), (2004)496–502.
- [8] Ankamwar, B., Chinmay, D., Absar, A., and Murali, S. Biosynthesis of gold and silver nanoparticles using *emblica officinalis* fruit extract, their phase transfer and transmetalation in an organic solution, *Journal of Nanoscience and Nanotechnology.*, 5(10), (2005) 1665–1671.
- [9] Shiv Shankar, S., Rai A., Ankamwar, B., Singh, A., Ahmad, A., and Sastry, M. Biological synthesis of triangular gold nanoprisms, *Nature Materials.*, 3(7), (2004)482–488.
- [10] Ankamwar, B. and Chaudhary, M. Gold nanotriangles biologically synthesized using tamarind leaf extract and potential application in vapor sensing, *Synthesis and Reactivity in Inorganic and Metal–Organic Chemistry.*, 35(1), (2005)19–26.
- [11] Laldhas, K.P., Cheriyan, V.T., Puliappadamba, V.T., Bava, S.V., Unnithan, R.G., Vijayammal, P.L., and

- Anto, R.J. A novel protein fraction from *Sesbaniagrandiflora* shows potential anticancer and chemopreventive efficacy, in vitro and in vivo, *J. Cell Mol. Med.*, 14(3), (2010) 636–646.
- [12] China, R., Mukherjee, S., Sen, S., Bose, S., Datta, S., Koley, H., Ghosh, S., and Dhar, P. Antimicrobial activity of *Sesbaniagrandiflora* flower polyphenol extracts on some pathogenic bacteria and growth stimulatory effect on the probiotic organism *Lactobacillus acidophilus*, *Microbiol. Res.*, 167(8), (2012)500–506.
- [13] Dethle, U.L., Joshi, S.S., Desai, S.S., and Aparadh, V.T. Screening of bioactive compounds of *Sesbaniagrandiflora* and *Pistiastratiotes*, *Indian Journal of Advances in plant Research (IJAPR)*, 1(1), (2014)27–30.
- [14] Devdatta, and Appanna, Nutritive value of Indian foods, *Indian Acad. Sci.*, (1954) 297–398.
- [15] Ramesh, P., Rajendran, A., and Meenakshisundaram, M. Green Synthesis of Zinc Oxide Nanoparticles Using Flower Extract *Cassia Auriculata*, *Journal of Nanoscience and Nanotechnology*, 2(1), (2014) 41–45.
- [16] Zhihong Jing., Lihua Tan., Fen Li., Jun Wang., Yucai Fu., and Qian Li. Photocatalytic and antibacterial activities of CdS nanoparticles prepared by solvothermal method. *Indian Journal of Chemistry*, 52A(1), (2013)57–62.
- [17] Lin, L., Wang, W., Haung, J., Li, Q., sun, D., Yang, X., Wang, H., He, N., Wang, Y., “Nature factory of silver nanowires: Plant-mediated synthesis using broth of *Cassia fistula* leaf,” *J. Chem. Eng.*, 162, (2010) 852–858.
- [18] S.R. Senthilkumar, T. Sivakumar, “*Green Tea (Camellia Sinensis) mediated synthesis of Zinc Oxide (ZnO) nanoparticles and studies on their antimicrobial activities*”, *International Journal of Pharmacy and Pharmaceutical Sciences*, ISSN- 0975-1491 Vol 6, Issue 6, (2014).
- [19] Song, J.Y., Jang, H.-K, Kim, B.S., “Biological synthesis of gold nanoparticles using *Magnolia kobus* and *Diopyros kaki* leaf extracts,” *Process Biochem.*, 44, (2009) 1133–1138.
- [20] Jain, N., Bhargava, A., Majumdar, S., Tarafdar, J.C., Panwar, J., “Extracellular biosynthesis and characterization of silver nanoparticles using *Aspergillus flavus* NJP08: a mechanism perspective,” *Nanoscale*, 3, (2011) 635–641.
- [21] Sagar Raut, Dr. P. V. Thorat, Rohini Thakre, “*Green Synthesis of Zinc Oxide (ZnO) Nanoparticles Using Ocimum Tenuiflorum Leaves*”, *International Journal of Science and Research (IJSR)*, Volume 4 Issue 5, May (2015).
- [22] K. Ravichandrika et.al, “*Synthesis, Characterization and Antibacterial Activity Of ZnO Nanoparticles*”, *International Journal Of Pharmacy And Pharmaceutical Sciences*, Vol 4, Issue 4, (2012) 336- 338.
- [23] Sudheesh K. Shukla, Eric S. Agorku, Hemant Mittal, Ajay K. Mishra, “Synthesis, characterization and photoluminescence properties of Ce³⁺ doped ZnO-nanophosphors”, *Chemical Papers*, (2012) DOI: 10.2478/s11696-013-0442-5.
- [24] Ravichandrika, P. Kiranmayi and R.V.S.S.N. Ravikumar, “*Synthesis, Characterization and antibacterial activity of ZnO Nanoparticles*”, *International Journal of Pharmacy and Pharmaceutical Sciences*, ISSN- 0975-1491 Vol 4, Issue 4, (2012).
- [25] R. Viswanatha, T.G. Venkatesh, C.C. Vidyasagar, Y. ArthobaNayaka, “Preparation and Characterization of ZnO and Mg-ZnO nanoparticles”, *Archives of Applied Science Research*, 4 (1)(2012) 480-486.
- [26] Elizabeth Varghese and Mary George, “Green Synthesis of Zinc Oxide Nanoparticles”, *International Journal of Advance Research in Science And Engineering IJARSE*, Vol. No.4, Issue No.01, ISSN-2319-8354(E), Jan (2015)307.
- [27] Davood Raoufi, *Journal of Luminescence*, 134, (2013)213-219.
- [28] R.Y. Hong, J.Z. Qian, J.X. Cao, *Powder Technol.* 163, (2006) 160.
- [29] R.Y. Hong, J.H. Li, L.L. Chen, D.Q. Liu, H.Z. Li, Y. Zheng, J. Ding, *Powder Technol.* 189, (2009) 426.
- [30] M. R. Alfaro Cruz, G. Ortega Zarzoza, G. A. MartínezCastañón and J. R. Martínez, “Thin films from different materials obtained by the Sol-Gel method: study of the morphology through Atomic Force Microscopy (AFM)”.